- CRADWICK, P. D. & HALL, D. (1971). Acta Cryst. B27, 1468.
- Das Sarma, B. & Bailar, J. C. (1955). J. Amer. Chem. Soc. 77, 5476.
- DODSON, G. G. & HALL, D. (1961). J. Inorg. Nucl. Chem. 23, 33.
- Fox, M. R. & LINGAFELTER, E. C. (1967). Acta Cryst. 22, 943.
- JEFFREY, G. A. & MCMULLAN, R. K. (1967). Progr. Inorg. Chem. 8, 43.
- LINGAFELTER, E. C. & BRAUN, R. L. (1966). J. Amer. Chem. Soc. 88, 2951.
- MAIN, P. & WOOLFSON, M. M. (1963). Acta Cryst. 16, 731. RAE, A. D. (1965). Acta Cryst. 19, 683.
- STEWART, J. M., LINGAFELTER, E. C. & BREAZEALE, J. D.

(1961). Acta Cryst. 14, 888.

Acta Cryst. (1972). B28, 49

less extensive.

The Crystal and Molecular Structure of 2-Dimethylaminoethyl Selenolbenzoate Hydrochloride

BY DAVID D. DEXTER

Department of Biochemistry, Columbia University, College of Physicians and Surgeons, 630 West 168th Street, New York, New York 10032, U.S.A.

(Received 18 November 1970)

The crystal structure of the local anesthetic 2-dimethylaminoethyl selenolbenzoate hydrochloride has been determined. The crystals are monoclinic, P_{21}/c with a=12.287 (3), b=6.945 (1), c=15.523 (3) Å and $\beta=94.57$ (2)°. Least-squares refinement on 2411 observed reflections resulted in an R value of

0.051. The N-C-C-Se group adopts the *trans* conformation with a torsion angle about the C-C bond of 174.4° . The two Se-C bond distances are both 1.945 Å with estimated standard deviations of 0.005 Å. The C-Se-C angle is 96.4 (2)°. Cell constants for four related compounds are also given.

Introduction

tices and 24 edges to the decahedron, but as basal faces

are shared there are 12 vertices and 20 edges per unit

of the column. There are then four hydrogen atoms

excess to each unit, one per three water molecules, and

these form hydrogen bonds to the phenolic oxygen

atoms of the nickel complex. The column of water mol-

ecules, infinite in one-dimension, represents an ex-

tremely compact assembly of water molecules provided

that hydrogen acceptors of appropriate number and

dimension are available, but it could not be extended

to a three-dimensional structure in their absence. The

structure is then in no sense a clathrate or potential

clathrate, but is an interesting intermediate between

these compounds and more conventional hydrates in

which the arrangement of water molecules, per se, is

The conformations of biologically active molecules related to acetylcholine (I) have been extensively studied (Chothia, 1970; Chothia & Pauling, 1970; Shefter & Mautner, 1969). These molecules all incorporate the

 $N-C-C-R_1$ group where R_1 may be oxygen, sulfur or selenium. While these short chains are formally flex-

ible, the N-C-C-O group usually adopts a *gauche* conformation (Pauling, 1968; Sundaralingam, 1968; Chothia & Pauling, 1970). Other studies show that the

N-C-C-S and N-C-C-Se groups take up the *trans* conformation (Shefter & Mautner, 1969).

In this work five members of a series of local anesthetics related to the tertiary analog of benzoylcholine, 2-dimethylaminoethyl benzoate (II, with $R_1 = O$, $R_2 = O$) were investigated. These compounds are obtained by systematically replacing the ether oxygen in the parent molecule (position R_1) with sulfur or selenium and the carboxyl oxygen (position R_2) with sulfur. Only the selenolester 2-dimethylaminoethyl selenolbenzoate hydrochloride (R_1 =Se, R_2 =O) was selected for complete structure determination.



Crystals of the parent compound $(R_1=0, R_2=0)$ proved highly deliquescent and were not investigated. The structure of the closely related 2-diethylaminoethyl *p*-aminobenzoate HCl (procaine HCl) was determined instead and a discussion of these results will appear subsequently (Dexter, 1972).



Experimental

Shown in Table 1 are the approximate cell dimensions of the five compounds as measured from Weissenberg and precession photographs. In all cases systematic absences h0l for l odd and 0k0 for k odd uniquely determine the space group as $P2_1/c$.

A crystal of the selenolester, approximately $0.18 \times 0.20 \times 0.14$ mm, was cut from a large, colorless plate and mounted with **b** along the φ axis of a General Electric computer-controlled diffractometer. The 2θ values (Cu $K\alpha_1$, $\lambda = 1.54050$ Å) of 15 carefully centered high angle reflections were used in a least-squares refinement of the lattice constants. (Crystal data are presented in Table 2.) The intensities of the 2810 independent reflections with $2\theta < 155^{\circ}$ (Cu $K\bar{\alpha}$, $\lambda = 1.5418$ Å) were measured using $\theta - 2\theta$ scans. The 399 reflections for which $I_{obs} < 2\sigma(I_{obs})$ were all treated as unobserved and assigned an intensity equal to 2σ (I_{obs}). The intensities were corrected for Lorentz and polarization factors but not for absorption.

During data collection the intensities of the three standard reflections decreased linearly at rates which seemed to be a function of 2θ . The crystal also became yellow-brown in color (possibly due to liberation of free selenium). Although the losses in intensity (9 to 33%) were significant, the solution of the structure proceeded uneventfully and exploratory refinement cycles led to a chemically reasonable structure.

Table 2.	Crystal	data f	or 2	2-dimet	hylami	noetl	iyl
	selenolb	enzoat	e hy	drochle	oride		

Formula	(CH ₃) ₂ N(CH ₂) ₂ Se COC ₆ H ₅ , HCl
M. W.	292.7
Space group	$P2_1/c$
a	12·387 (3) Å
Ь	6.945(1)
с	15.523(3)
β	94·57(2)°
Volume	1331·2 Å ³
ρ(X-ray)	1.460 g.cm ⁻³
ϱ (flotation)	1.47 g.cm^{-3}
Ζ	4
μ	59.3 cm^{-1}

The estimated standard deviations of the lattice constants are three times those obtained from the least-squares calculation.

Structure determination and refinement

The selenium atom and chloride ion were readily located from an E^2-1 vector map. Coordinates and

Table 3. Final positional and anisotropic thermal parameters for the non-hydrogen atoms

The positional parameters for Cl⁻ and Se are × 10⁵. All others are × 10⁴. The estimated standard deviations are those obtained from the least-squares and refer to the last decimal place given. The form of the anisotropic temperature factor is exp $[-\frac{1}{4}(B_{11}h^2a^{*2}+B_{22}k^2b^{*2}+B_{33}l^2c^{*2}+2B_{12}hka^*b^*+2B_{13}hla^*c^*+2B_{23}klb^*c^*)].$

	x	У	Z	B ₁₁	B ₂₂	B33	B_{12}	B_{13}	B_{23}
C(1)	3598 (4)	1948 (6)	7904 (3)	4.9 (2)	4.4 (2)	3.1 (1)	-0.2(1)	-0.7(2)	0.1 (1)
C(2)	5371 (3)	2051 (7)	8721 (3)	3.8 (1)	4.6 (2)	5.0 (2)	0 ∙3 (1)	0.7 (2)	0·4 (1)
C(3)	3744 (2)	1892 (4)	9518 (2)	3.7 (1)	2.7(1)	3.0 (1)	0.0 (1)	-0.2(1)	0·6 (1)
C(4)	2635 (3)	2699 (6)	9633 (2)	3.4 (1)	4.2 (2)	3.7 (1)	0.4(1)	0.2(1)	1.0 (1)
C(5)	8458 (2)	4451 (5)	4650 (2)	3.4 (1)	4.5 (2)	2.9 (1)	0.3 (1)	0.2(1)	0.2 (1)
C(6)	8747 (2)	3128 (5)	3950 (2)	3.0 (1)	3.9(1)	2.9 (1)	-0.3(1)	0.2(1)	0.1 (1)
C(7)	8210 (3)	3159 (5)	3128 (2)	3.0 (1)	4.4 (2)	$3 \cdot 1$ (1)	-0.0(1)	-0.1(1)	0.0 (1)
C(8)	8510 (3)	1892 (5)	2500 (2)	3.7 (1)	5.1 (2)	3.4 (1)	-0.8(1)	-0.1(1)	-0.7(1)
C(9)	9353 (3)	0642 (6)	2680 (2)	4.6 (2)	4.5 (2)	4.5 (2)	-0·1 (1)	0.5 (2)	-1.0(1)
C(10)	9894 (3)	0580 (6)	3504 (3)	4·2 (1)	4.4 (2)	5.2 (2)	0.8 (1)	-0.0(2)	-0.3(1)
C(11)	9587 (3)	1804 (5)	4133 (2)	4·1 (1)	4.3 (2)	3.7 (2)	0.2(1)	-0.6(1)	0.2(1)
N	4207 (2)	2611 (4)	8711 (2)	3.5 (1)	2.4 (1)	2·9 (1)	-0.1(1)	-0.1(1)	0.4 (1)
С	8635 (3)	4106 (5)	5406 (2)	7.3 (2)	6.2 (2)	2.6 (2)	1.6 (1)	0.1 (1)	0.1 (1)
C1-	58321 (8)	19387 (1)	62870 (6)	5.35 (4)	2.22 (3)	4.45 (4)	-0·08 (2)	-0.62(3)	-0.02(2)
Se	21848 (3)	18784 (6)	07473 (2)	4.98 (2)	4.76 (2)	3.64(2)	-0.51(1)	0.47(2)	-0.25(1)

H(1)

H(2) H(3) H(4) H(5) H(6) H(7) H(8) H(10) H(11) H(12) H(13) H(14) H(15)

H(16)

isotropic thermal parameters for the heavy atoms and the scale factor were adjusted with one cycle of fullmatrix least squares (with unit weights) to give an Rvalue of 0.35. The subsequent electron density map revealed the other 13 non-hydrogen atoms. After several least-squares cycles with individual isotropic temperature factors R decreased to 0.12. At this point the hydrogen atoms were located from a difference map and included in the refinement with fixed isotropic temperature factors. After allowing the heavy atoms to refine anisotropically R was 0.075. Weights were then assigned from a plot of $|\Delta F|$ vs. $|F_o|$. Several intense low angle reflections appeared to be affected by extinction and were excluded from the remainder of the refinement. Using anisotropic thermal parameters for all non-hydrogen atoms and isotropic thermal parameters for hydrogen atoms, the refinement converged with R = 0.051. In the last cycle for the 200 parameters refined, the average change was 0.04σ and the maximum change was 0.4σ . The final difference map showed a peak of about 0.5 electron at the position of the chloride ion and a pattern of peaks and



Fig. 1.(a) Bond distances in 2-dimethylaminoethyl selenolbenzoate hydrochloride. (b) Bond angles involving non-hydrogen atoms. (Angles involving H atoms are within expected limits.)

holes about the selenium atom position with heights of about ± 1.0 electron. While this pattern might suggest incomplete refinement of the anisotropic thermal parameters, it also could be a manifestation of the crystal decay noted during the data collection.

Scattering factors for carbon, nitrogen and oxygen were taken from *International Tables for X-ray Crystallography* (1962), for selenium and chlorine (Cl⁻) from Cromer & Waber (1964) and for hydrogen from Stewart, Davidson & Simpson (1965). The bulk of the computations was carried out with the *X-ray* 67 *System* programs (Stewart, 1967). The mean plane program used was that of Ahmed, Hall, Pippy & Saunderson (1966).

Table 4. Final coordinates and isotropic thermal parameters for the hydrogen atoms

The positional parameters are all $\times 10^3$.

x	У	Z	В
356 (5)	050 (9)	792 (4)	8 (1)
379 (4)	239 (8)	736 (4)	7 (1)
309 (3)	217 (6)	787 (3)	3 (1)
545 (4)	070 (9)	866 (3)	6 (1)
574 (4)	262 (7)	919 (3)	5 (1)
560 (4)	245 (8)	830 (3)	5 (1)
377 (3)	061 (6)	957 (2)	3 (1)
420 (2)	217 (5)	995 (2)	3 (1)
205 (5)	242 (9)	926 (4)	7 (1)
269 (3)	405 (7)	962 (3)	4 (1)
761 (3)	398 (6)	303 (3)	4 (1)
825 (4)	194 (7)	198 (3)	6 (1)
956 (4)	-018 (7)	225 (3)	6 (1)
046 (4)	-018(8)	369 (3)	7 (1)
002 (4)	188 (6)	465 (3)	6 (1)
418(3)	410 (6)	871 (2)	3 (1)

Table 5. Torsion angles

	Angle (τ)
C(11)-C(6)-C(5)-Se	153·2°
C(6) - C(5) - Se - C(4)	169.4
O - C(5) - Se - C(4)	166-9
C(5)—Se— $C(4)$ – $C(3)$	87.4
Se— $C(4)-C(3)-N$	174.4
C(4) - C(3) - N - C(2)	169.7
C(4) - C(3) - N - C(1)	65.9



Fig. 2. Stereo diagrams (Johnson, 1965) showing molecular packing along b. $a \rightarrow, c \uparrow$. Hydrogen atoms have been omitted. Chloride ions have been shaded.

52

Results and discussion

Final coordinates and thermal parameters for nonhydrogen atoms are given in Table 3 and for hydrogen atoms in Table 4. Fig. 1 shows bond distances, bond angles and the numbering system. Some torsion angles along the side chain are listed in Table 5. Table 6 is the list of observed and calculated structure factors.

Table 6. Observed and calculated structure factors

The columns are l, $10F_o$ and $10F_c$. Unobserved reflections are marked with an asterisk. Reflections marked with an E were excluded from the final refinement cycles.

Table 6 (cont.)

-100 97 32 15 -657 -157 -157 -157 -107 -107 50 -25 -28 -190 73 -74 09001 ... 92112799883397380001411278 ... 2019783131898207122847350 ... 22311144997098826372354441022171558 ... 2543362228737 440,72001477471177476 . 2013720773331777811457641057801712217 10 22/3000003540077807812700072400071407710 . 1 7200734530007240172101 104 115 12 -13 -1 -2 -3 -4 -5 -6 -7 -8 -9 -100 -111 -112 -13 -14 054 + 76107151A #607353558 #657387 #32 8884 +54+38814×9833507841552379 + 5 47515356 #51757348 #613725342,919 109476878978978 #5 109476878978978 #5 109476878978978 #5 109476878978978 #5 109476878978978 #5 109476878978978 #5 109476878978978 #5 109476878978978 #5 109476878978978978 #5 109476878978978 #5 109476878978978 #5 109476878978978 #5 109476878978978 #5 109476878978978 #5 109476878978978 #5 109476878978978 #5 109476878978978 #5 109476878978978 #5 109476878978978 #5 109476878978978978 #5 109476878978978978878 #5 109476878978978 #5 109476878978978978978 #5 109476878978978978978978978978978978978978 244240022771133828 67, 760544152925627476233626909 7, 76060454474850070530478 7, 2528221322271547578 7, 4729216542 4, 455777231482114 -21 -14 170 128 106 122 32 87 -37 -9 -9 -10 -11 -15 6-778899001112234 -9 117 49 186 -42 63 -129 63 -129 63 -38 -77 37 -1 -2 -3 -4 - -5 -6 - -7 -8 8 -9 9 -10 -11 12 -13 -13 -12 -13 -77 -8 -99 -10 -11 011423344359067788 114233445598778 011422334455967 173 45 22+ 37 180 23+ 34 6+++L 3655 2771 1076 3172 -1230 -1200 -1000 -100 -1000 178 91 34 198 -5 32 -13 -14 -14 -15 -15 -15 ??? 124 - 075 0-1-2-3-4 4 5 5 6 6 7 7 8 9 011223344556677810 20 -188 -30 -110 -117 -117 -127 -38 -3 -3 -3 78 117 52 93 78 -127 -32 -0 -129 -124 -52 60 -52 -32 0-1-22334455067789 1458652210124661555433223110287045604220229 + 0 0542387872241220248862673308862073048 + 7116000081660893171 0112213445566778899001111223344556 4] -? 8 1234334 -22422 -2242 -2242 -8037 - $\begin{array}{c} -102\\ 5\times7\\ -1406\\ 6& 6& 1\\ 1& 1& 2& 6\\ 1& 1& 1& 2& 6\\ 1& 1& 2& 2& 7\\ 1& 1& 2& 2& 6\\ 1& 1& 2& 2& 7\\ 1& 1& 5& 7& 7\\ -& 1& 1& 5& 7\\ 1& 1& 5& 7& 7\\ -& 1& 1& 5& 7\\ 1& 1& 1& 5& 7\\ 1& 1& 1& 5& 7\\ 1& 1& 1& 5& 7\\ 1& 1& 1& 5& 7\\ 1& 1& 1& 1& 7\\ 1& 1& 1& 1& 7\\ 1& 1& 1& 1& 7\\ 1& 1& 1& 1& 7\\ 1& 1& 1& 1& 7\\ 1& 1& 1& 1& 7\\ 1& 1& 1& 1& 7\\ 1& 1& 1& 1& 7\\ 1& 1& 1& 1& 1\\ 1& 1& 1& 1& 7\\ 1& 1& 1& 1& 1\\ 1& 1& 1& 1\\ 1& 1& 1&$ -1-223344567 -46 -147 -31 -34 -1 -74 27 1 37 -10 -6 117 27 -93 -102 104 -38 83 -8 107 -101 94 27 38 0112233445566775899101 24 77 121 2 44 -51 -89 148 -39 -39 24 -39 -39 24 -39 -12 65 32 189 -122 -765 -1051 -582 -420 0123456783011234 12345679931123450 011423544556677788993311122354 C-11223344555607788990001111233445 -82 -62 -28 -13 -64 -54 71 -49 (1122)34455607788590[11123 2270387 -038715 -0751-445 -02751-445 -003309 -0300308 12803309 19 - 25 - 22 - 22 - 18 - 40 - 40 - 40 -1340 1360 1560 1560 1560 1560 1760 -1062 1076 -1062 1076 -1076 -10752 C 1 1 2 2 3 3 4 4 5 4 6 6 7 7 8 8 0 6 C 1 1 2 2 3 3 4 4 5 4 6 6 7 7 8 8 0 6 C 1 1 2 2 3 3 4 4 71122334455667788997111 71122334455667788997101 111 -31 -211 -111 -137 -22 125 155 -55 1 2 3 5 6 7 8 9 10 11 0 1 -1 2 2 3 3 4 4 5 5 6 6 7 7 88 8 9 9 0 100 -111 122 3 13 -66 -137 -276 -117 178 -65 787 293 13 -255 -66-201 -14 -------174 -31 -14 -85 -21 7 -8 16 170 -6 0-1-2233445566778 17 133 -45 70 154 -18 41 13 -140 -30 -140 -158 -370 -25 -61 127 51 120 40 16 -67 -180 -107 -62 7 -175 612 143 75 -13 01122004405 13 -50 167 -210 -107 -31 -7 -130 57 33 -27 64 -13 203 r 164 1 183 -171 0

As expected, the N-C-C-Se group in 2-dimethylaminoethyl selenolbenzoate hydrochloride adopts the *trans* conformation. The torsion angle about the C(3)-C(4) bond is 174.4°. In structures of related selenol- and thiolesters the corresponding torsion angles range from 171 to 176° (Shefter & Mautner, 1969). Karle & Estlin (1969) found that the conformation of

the N–C–C–Se chain in 2-aminoethylselenic acid HCl is also *trans* with a torsion angle of 171.6° . It is very probable that the *trans* conformation is maintained in the four nearly isomorphous crystals (Table 1) investigated here.

The two chemically different C-Se distances are both 1.945 (5) Å and the C(5)-Se-C(4) angle is 96.4 (2)°. These distances are in good agreement with the others which have been tabulated by Hope, Knobler & McCullough (1970). They determined a weighted average Se-C distance for nonaromatic compounds of 1.98 (2) Å and 1.92 (2) Å for aromatic compounds. A comparison of the selenolester group found here with that reported for acetylselenolcholine iodide (Shefter & Kennard, 1966) would not be very meaningful because of the large standard deviations quoted for the latter structure.

The benzene ring is significantly nonplanar. For

the mean plane of all six atoms $C(6) \cdots C(11)$, 0.6616X +0.6922Y-0.2882Z-1.497=0, χ^2 =300. (The coefficients refer to an orthogonal set of axes where X lies along **a**, Y is in the (**a**,**b**) plane and Z lies along **c***.) χ^2 =2.6 for the mean plane 0.6682X-0.6887Y-0.2811Z-1.4241=0 through the four most nearly coplanar atoms $C(8) \cdots C(11)$. C(6) is 0.035 (3) Å below and C(7) 0.052 (4) Å above this plane.

There is little evidence of resonance interaction between the selenolester and phenyl groups. The C(5)– C(6) bond distance of 1.489 (5) Å is only 3.8σ shorter than the other C-C single bond in the molecule. The angle between the plane of the selenolester group O-C(5)-Se, 0.9163X - 0.4000Y - 0.0166Z - 1.8542 =0, and the mean plane of the whole phenyl ring is 25.4°. This twist places the Se atom on the opposite side of the mean plane from C(7).

The conformation about the Se–C(4) bond ($\tau = 87 \cdot 4^{\circ}$) is not that normally found in esters (τ about 180°) (Mathieson, 1965) or in other selenol- and thiolesters (τ about 120°) (Shefter & Mautner, 1969). A similar conformation ($\tau = 79^{\circ}$) has been found in acetylcholine bromide (Canepa, Pauling & Sörum, 1966). Liquori, Damiani & deCoen (1968) concluded on the basis of total van der Waals energy calculations that in acetylcholine the observed gauche conformation was not significantly less stable than the trans. In the present structure, the observed torsion angle probably arises to accommodate the formation of different environments for the alkyl and aromatic parts of the molecule. A packing diagram (Fig. 2) shows that the protonated nitrogen atoms are hydrogen bonded to chloride ions (N-Cl, 3.07 Å) so that layers of charged groups alternate with layers of interlocking phenyl rings.

54

The smoothness of the refinement and the reasonableness of the final structure prompted an attempt to assess the effect of the crystal decay on the results.

The hk0 reflections were measured twice: at the beginning of the experiment as part of the hkl reflections (I_o) and later as part of the $hk\bar{l}$ reflections $(I_{1/2})$. Fig. 3 is a plot of $\ln(I_{1/2}/I_0)$ versus $\sin^2\theta/\lambda^2$. Assuming that the decrease in intensity can be described in terms of an apparent B parameter, by the end of the data collection the total effect of the decomposition was an increase in the overall temperature factor by about 1.5 Ų.

The effect of a larger temperature factor is to increase the uncertainty in atomic positions, not to displace them (Cruickshank, 1949). It seems reasonable, therefore, to accept as valid the positional parameters ob-



Fig.3. Plot of $\ln(I_{1/2}/I_0)$ versus $(\sin^2\theta)/\lambda^2$ showing loss of intensity at midpoint of data collection for 75 hk0 reflections. The slope of the least-squares straight line corresponds to a B parameter of 0.8 (1) Å².

tained from the data affected by crystal decay - with the usual warning that the standard deviations obtained from the least-squares probably will be too small.

The author wishes to thank Professor Barbara W. Low, who suggested this problem, for her critical interest and advice. Professor H. G. Mautner kindly provided the crystals used in this investigation. This work was supported by the National Institutes of Health Grant NS 07747 and in part by the National Science Foundation Grant GB 7272.

References

- AHMED, F. R., HALL, S. R., PIPPY, M. E. & SAUNDERSON, C. P. (1966). NRC Crystallographic Programs for the IBM/360 System. IUC World List of Crystallographic Computer Programs. 2nd ed. Appendix, p. 52.
- CANEPA, F. G., PAULING, P. J. & SÖRUM, H. (1966). Nature, Lond. 210, 907.
- CHOTHIA, C. H. (1970). *Nature, Lond.* **225**, 36. CHOTHIA, C. H. & PAULING, P. (1970). *Proc. Nat. Acad.* Sci. 65, 477.
- CROMER, D. T. & WABER, J. T. (1965). Acta Cryst. 18, 104.
- CRUICKSHANK, D. W. J. (1949). Acta Cryst. 2, 154.
- DEXTER, D. D. (1972). Acta Cryst. B28, 77.
- HOPE, H., KNOBLER, C. & MCCULLOUGH, J. D. (1970). Acta Cryst. B26, 628.
- International Tables for X-ray Crystallography (1962). Vol. III. Birmingham: Kynoch Press.
- JOHNSON, C. K. (1965). A Fortran Ellipsoid Plot Program for Crystal Structure Illustrations. ORNL-3794, Oak Ridge National Laboratory, Oak Ridge, Tennessee.
- KARLE, I. L. & ESTLIN, J. A. (1969). Z. Kristallogr. 129, 147.
- LIQUORI, A. M., DAMIANI, A. & DECOEN, J. L. (1968). J. Mol. Biol. 33, 445.
- MATHIESON, A. MCL. (1965). Tetrahedron Letters, p. 4137.
- PAULING, P. (1968). In Structural Chemistry and Molecular Biology. Edited by A. RICH and N. DAVIDSON, p. 555. San Francisco, London: Freeman.
- SHEFTER, E. & KENNARD, O. (1966). Science, 153, 1389.
- SHEFTER, E. & MAUTNER, H. G. (1969). Proc. Nat. Acad. Sci. 63, 1253.
- STEWART, J. M. (1967). Technical Report 67-58, Computer Science Center, University of Maryland.
- STEWART, R. F., DAVIDSON, R. E. & SIMPSON, W. T. (1965). J. Chem. Phys. 42, 3175.
- SUNDARALINGAM, M. (1968). Nature, Lond. 217, 35.